CO₂ Capture by Nickel Hydroxide Interstratified in the Nanolayered Space of a Synthetic Clay Mineral

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ABSTRACT: Clay minerals can adsorb large amounts of CO₂ and are present in anthropogenic storage sites for CO₂. Nanoscale functionalization of smectite clay minerals is essential for developing technologies for carbon sequestration based on these materials and for safe-guarding relevant long-term carbon storage sites. We investigate the adsorption mechanisms of CO₂ in dried and hydrated synthetic Ni-exchanged fluorohectorite clay—using a combination of powder X-ray diffraction, Raman spectroscopy, and inelastic neutron scattering. Both dried and hydrated Ni-exchanged fluorohectorite show crystalline swelling and spectroscopic changes in response to CO₂ exposure. These changes can be attributed to interactions with [Ni(OH)0.83(H₂O)1.17]0.37-interlayer species, and swelling occurs solely in the interlayers where this condensed species is present. The experimental conclusions are supported by density functional theory simulations. This work demonstrates a hitherto overlooked important mechanism, where a hydrogens species present in the nanospace of a clay mineral creates sorption sites for CO₂.

Continuous and increasing release of greenhouse gases, in particular CO₂ to the atmosphere has a detrimental impact on our environment. In addition to emission reduction, it is imperative to establish effective ways of capturing these gases. Clay minerals have become an attractive alternative for investigating CO₂ capture because of their excellent behavior in adsorption and catalysis. This is in addition to their abundance and stability, which make them potentially scalable for industrial processes.

Smectite clay minerals, such as hectorite and montmorillonite, are known to readily swell in response to water, and organic compounds. They consist of two-dimensional nanolayered stacks possessing a negative layer charge due to nonequivalent substitutions of atoms in the layers. The negative layer charge is typically counterbalanced by exchangeable interlayer cations, for example Na⁺, Ca²⁺, and K⁺.

The sorption of CO₂ in smectite clays at various hydration states as well as under gaseous and supercritical CO₂ conditions has been studied by a wide range of techniques, such as X-ray and neutron diffraction, infrared spectroscopy, nuclear magnetic resonance, quartz crystal microbalance, and volumetric/gravimetric gas adsorption. The sorption of CO₂ in smectite clays has been shown to largely depend on the initial hydration, the interlayer cation, and the specific clay type. For montmorillonite (Mt) and OH-hectorite (Hec), experiments and simulations at 323 K and 90 bar have shown that for smaller interlayer cations with a high hydration energy (e.g. Na⁺, Ca²⁺, and Mg²⁺), a submonolayer of water is necessary to open up the interlayer to swell in response to CO₂. However, for larger interlayer cations with low hydration energy (e.g. Cs⁺, K⁺, and NH₄⁺), CO₂ is intercalated even in the absence of water.

Natural Hec has a varying composition of F⁻ and OH⁻ groups, whereas fluorohectorite is a synthetic Hec with only fluorine groups. Simulations have shown that by varying the F⁻/(F⁻ + OH⁻) composition, the CO₂/(CO₂ + H₂O) ratio increases with increasing F⁻ substitution. It has previously been demonstrated that a fluorinated Hec exchanged with Na⁺, Li⁺, and Ni²⁺ swells in response to CO₂ exposure and that Ni-Hec showed a much larger adsorption of CO₂ than Li-Hec and Na-Hec.

Here, we report a hitherto overlooked mechanism for CO₂ sorption in Ni-Hec as studied with synchrotron X-ray diffraction, Raman spectroscopy, and inelastic neutron scattering (INS). The intercalation properties of Ni²⁺ in smectites have been shown to be substantially different from
other interlayer cations.\textsuperscript{5,23} As recently reported in Loch et al.,\textsuperscript{24} Ni-Hec forms an ordered interstratification with a chloride-like condensed [Ni(OH)\textsubscript{0.63}(H\textsubscript{2}O)\textsubscript{1.17}]\textsuperscript{0.37} species in one interlayer and a smectite-like structure with hydrated Ni\textsuperscript{2+} cations in the adjacent interlayer. By employing powder X-ray diffraction (PXRD), we investigate the structural changes in response to CO\textsubscript{2}. We utilize \textit{in situ} Raman scattering combined with PXRD, which allows studying how CO\textsubscript{2} adsorption changes the local structure. INS has no selection rules for the vibrational modes and a large cross section of the hydrogen atoms, making the signal from the water motions and other hydrogenous species, such as nickel hydroxide, prominent.\textsuperscript{25} INS has previously been successfully employed to study how layer charge and cations influence the hydration properties of saponite, montmorillonite, and beidelitte.\textsuperscript{26,27}

By combining PXRD, Raman spectroscopy, and INS, we elucidate the active mechanisms of CO\textsubscript{2} sorption for Ni-Hec in the dried (at 150 °C under dynamic vacuum) and hydrated (equilibrated at 43% r.h.) states. To support our experimental conclusions, we have performed density functional theory calculations (DFT).

### METHODS

#### Materials

Na-Hec with a stoichiometric composition of Na\textsubscript{(Mg, Li)}\textsubscript{4} Si\textsubscript{4} O\textsubscript{10} F\textsubscript{2} was prepared via melt synthesis according to a published procedure,\textsuperscript{7} followed by annealing (6 weeks, 1045 °C) to improve charge homogeneity and phase purity.\textsuperscript{28} Ni-Hec was prepared by cation exchange of Na-Hec with 0.2 M nickel−acetate solution (>10 fold excess of the CEC, 5 times). The exchanged Ni-Hec was washed 5 times with Millipore water. The hydration and drying procedures are described below for each type of experiment.

#### In Situ Powder X-ray Diffraction and In Situ Raman Spectroscopy

For Ni-Hec, combined Raman and synchrotron X-ray powder diffraction data were collected at the Swiss-Norwegian Beamlines (SNBL, station BM01B) at ESRF, Grenoble, France. PXRD data were measured at a wavelength of 0.77495 Å using a 2D PILATUS2M detector.\textsuperscript{29} The 2D diffractograms were integrated using the SNBL Bubble software package.\textsuperscript{30} The 2D diffraction patterns. The powdered samples were contained in borosilicate capillaries mounted in a custom made high-pressure sample cell based on the design by Jensen et al.\textsuperscript{31} The temperature of the capillary was controlled by an Oxford Cryostream 700+ nitrogen blower. The cell was connected to a gas handling system providing vacuum (10\textsuperscript{-6} mbar) from a turbomolecular pump or pressurized CO\textsubscript{2} of quality 99.9995% for 20 min, and subsequently cooling down to T < 20 K with CO\textsubscript{2} still inside the sample holder. Each INS measurement lasted around 6–8 h, depending on the amount of hydrogen in the sample.

The data were reduced using the software package Mantid.\textsuperscript{34}

#### DFT Simulations

We performed \textit{ab initio} investigations based on density functional theory calculations\textsuperscript{35,36} considering the generalized gradient approximation with the state-of-the-art van der Waals-corrected density functional developed by Berland and Hyldgaard,\textsuperscript{37} as implemented in the Siesta package.\textsuperscript{38} Norm-conserving pseudopotentials were used with double-zeta plus polarization and spin-polarized localized atomic orbital basis sets considering an energy cutoff of 400 Ry. The crystallographic structure reported by Breu et al.\textsuperscript{39} was used for the Ni-Hec model with the unit cell containing 39 atoms. For β-Ni(OH)\textsubscript{2}, we took the structural model from the Aflowlib repository.\textsuperscript{40} Adsorbents were allowed to relax during structural optimization using the conjugate gradient method under the self-consistent cycles considering a criteria of 0.01 eV/Å for the force minimization. The adsorption energy \(E_{\text{Ads}}\) was calculated following the equation

\[
E_{\text{Ads}} = E_{\text{Total}} - E_{\text{Crystal}} - E_{\text{Molecule}}
\]

where \(E_{\text{Total}}\) is the total energy for the adsorbed molecule in the most favorable adhesion site; \(E_{\text{Crystal}}\) and \(E_{\text{Molecule}}\) are the total energies for the isolated systems.

### RESULTS AND DISCUSSION

#### Powder X-ray Diffraction. Dried Ni-Hec. In Figure 1, the swelling of Ni-Hec at 300 K with increasing CO\textsubscript{2} pressure is shown. After drying the sample in vacuum at 150 °C, an intense Bragg reflection is observed at 0.551 Å\textsuperscript{-1}, corresponding to a d-value of 11.40 Å, which recently was identified as the...
Hydrated Ni-Hec. In Figure 2, the expansion upon CO₂ exposure of a Ni-Hec hydrated at 300 K is shown, adopting the 2 WL hydrate state with a (002) Bragg reflection at 14.82 Å,24 which gradually increases to 15.06 Å as the CO₂ pressure increases. Compared to the dried sample, a smaller increase is observed, that is 0.48 versus 1.78 Å. Again, the width of the Bragg reflection at the initial state and final CO₂ exposed state is comparable.

Interestingly, upon swelling in response to CO₂, the crystallinity was improved by reducing planar defects as indicated by cross-reflections in the region of the 11/02 band gaining significant intensity (Figure 3). While the starting material shows a \( \lambda \)-shaped band typical for turbostratic stacking, after CO₂ saturation, a 3D ordered stacking is observed with a refined unit cell of \( (C_2/m, a = 5.2504 \text{ Å}, b = 9.0780 \text{ Å}, c = 30.25 \text{ Å}, \beta = 96.6280^\circ, \alpha/\gamma = 90^\circ) \). Such a disorder/order transformation has previously been observed at the transition from 1 WL to 2 WL hydrate of Na-Hec4,28 and has been attributed to a strengthened interaction of interlayer species with the surface of the 2:1-silicate layers via hydrogen bonding. In the present case, this disorder/order transformation was reversible and was lost again as the CO₂ pressure was reduced, as shown in Figure S5.

In Situ Raman Spectroscopy. In situ Raman spectra corresponding to pristine (not yet exposed to CO₂) hydrated Ni-Hec (Figures 2 and 3) and exposed to 40 bar CO₂ pressure are shown in Figure 4. The assignment of the vibrations is given in Table 1. A very high fluorescent background did not allow us to record Raman spectra for dried Ni-Hec. Fluorescence is commonly observed in clay minerals,42 however, in this study, fluorescence also arises from the electronic structure of the Ni²⁺ cations, which have a 3d⁸ valence configuration.43 Local distortion of octahedrally coordinated chlorite-type \([\text{Ni(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17} \) greatly affects the associated orbitals also contributing to the observed fluorescence.43

Modes associated to the Hec structure are assigned according to Rinaudo et al.24 As recently shown in Loch et al.,24 Ni-Hec contains both Ni²⁺ and \([\text{Ni(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17} \) species segregated into adjacent interlayers. Thus, the vibrations at 319 cm⁻¹ (#4) and 467 cm⁻¹ (#7) can be associated to Ni–OH vibrations.35 We also observe a third feature around 506 cm⁻¹ (#8) that can be
assigned either to the presence of structural defects and/or to a \( E_g(R) \) mode for the Ni–OH structure.

In the range of OH-modes between 3000 and 3600 cm\(^{-1}\), the sharp feature at 3565 cm\(^{-1}\) (\#17) is assigned to the \( A_{1g} \) \( O^–H \) stretch of Ni–OH.\(^{45}\) The remaining modes in this region are assigned to water present in the clay. These can be divided into water bonded directly to the interlayer cations, interlayer water that does not directly interact with the cations, and water outside the interlayers.\(^{28}\) The latter constitutes a tiny amount because of the large aspect ratio of the clay particles,\(^{28}\) which results in a relatively low external surface area. The water not directly coordinated to the interlayer cations can either be in the chlorite-like or the smectite-like layers. Given these possibilities, an unambiguous assignment of the modes in this region is difficult.

After exposure to 40 bar of CO\(_2\) (Figure 4), the intensity of bands related to CO\(_2\) at 1286 cm\(^{-1}\) (\#14) and 1388 cm\(^{-1}\) (\#15) increases. There is also a clear change in the intensity associated to OH-groups between 3000 and 3600 cm\(^{-1}\) and an increase of intensity and sharpening of the vibration associated to the Mg, LiO\(_6\) octahedra, and the asymmetric stretch of SiO\(_4\). Because the density ratio between intercalated and non-intercalated gaseous CO\(_2\) in the probing volume of the Raman laser is low, the intensity of the CO\(_2\) vibrations is most likely to be dominated by gaseous CO\(_2\) in the capillary. The increase and sharpening of the modes associated to the octahedral and the tetrahedral sheet of the clay layers can be associated to the order−disorder transition observed by PXRD (Figures 2 and 3). Indeed, an increased long-range order in clays is known to enhance the signal to noise-ratio and sharpen the peaks.\(^{47}\) The shape change in the H\(_2\)O/OH-region could indicate some local modification of the water population in the interlayers. In

![Figure 3. X-ray diffraction pattern of hydrated Ni-Hec as a function of CO\(_2\) pressure at 300 K.](https://dx.doi.org/10.1021/acs.jpcc.0c07206)

![Figure 4. Raman spectra of hydrated Ni-Hec as prepared at 1 bar and exposed to 40 bar of CO\(_2\) at 300 K. Numbering indicates the assigned vibrations, and arrows highlight prominent changes in the spectra. Mode assignment is summarized in Table 1.](https://dx.doi.org/10.1021/acs.jpcc.0c07206)

Table 1. Raman Vibration Assignment of the Peaks

<table>
<thead>
<tr>
<th>feature number</th>
<th>position (cm(^{-1}))</th>
<th>suggested assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>184</td>
<td>MgLiO(_6) octahedra(^{44})</td>
</tr>
<tr>
<td>2</td>
<td>222</td>
<td>(small vibration) unassigned</td>
</tr>
<tr>
<td>3</td>
<td>280</td>
<td>O–H–O vibrations(^{45})</td>
</tr>
<tr>
<td>4</td>
<td>319</td>
<td>( A_{1g}(T) ), ( E_g(T) ) or ( E_g(R) ) Ni–(OH)(^{45})</td>
</tr>
<tr>
<td>5</td>
<td>330</td>
<td>bending SiO(_4)(^{44})</td>
</tr>
<tr>
<td>6</td>
<td>371</td>
<td>stretch Mg–O(^{19})</td>
</tr>
<tr>
<td>7</td>
<td>467</td>
<td>( A_{1g}/E_g(LO)/A_{2u}(LO) ) Ni–(OH)(^{45})</td>
</tr>
<tr>
<td>8</td>
<td>506</td>
<td>( A_{1g}(T) ) or ( E_g(T) ) Ni–(OH)(^{45})</td>
</tr>
<tr>
<td>9</td>
<td>696</td>
<td>( \nu_5 ) SiO(_4)(^{44})</td>
</tr>
<tr>
<td>10</td>
<td>783</td>
<td>vibration SiO(_4)(^{20})</td>
</tr>
<tr>
<td>11</td>
<td>1006</td>
<td>vibration SiO(_4)(^{20})</td>
</tr>
<tr>
<td>12</td>
<td>1105</td>
<td>( \nu_6 ) SiO(_4)(^{44})</td>
</tr>
<tr>
<td>13</td>
<td>1265</td>
<td>CO(_2) hot band(^{11})</td>
</tr>
<tr>
<td>14</td>
<td>1286</td>
<td>CO(_2) 2( \nu_2 )(^{51})</td>
</tr>
<tr>
<td>15</td>
<td>1388</td>
<td>CO(_2) ( \nu_1 )(^{21})</td>
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<td>17</td>
<td>3565</td>
<td>Ni–(OH) ( A_{1g} ) O–H(^{45})</td>
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</table>
addition, there are clear changes for the OH-stretching bands. The observed changes for the OH-stretching vibrations bear resemblance to previously observed changes when β-Ni(OH)₂ transforms to α-Ni(OH)₂,⁴⁸ which indicates that CO₂ has an effect on the intercalated [Ni(OH)₀.₈₅(H₂O)₁.₁₇]₀.₃₇¹.₁₇⁺ species. The observed increase of intensity around 1100 cm⁻¹ (#12) can then, in addition to ordering of the clay crystal structure, be associated with introduced stress in the nickel hydroxide lattice.⁴⁸ The increased intensity at 1100 cm⁻¹ (#12) could also be due to carbonate (CO₂⁻⁵), which is known to form a strong symmetric stretch υ₁ mode in this region for Raman spectroscopy.⁴⁹ The remaining modes for carbonate, which would be located around 880, 1400, and 680 cm⁻¹, are weaker⁴⁹ and, in this case, located in regions where other modes due to the clay and CO₂ gas phase are dominating.

**Inelastic Neutron Spectroscopy.** INS spectra of pristine Ni-Hec, both in dried and equilibrated at 43% r.h. states, are presented in Figure 5.

![INS spectra of Ni-Hec in dried and hydrated states.](image)

**Figure 5.** INS experimental spectra of Ni-Hec in dried (145 °C under vacuum) and hydrated states (equilibrated at 43% r.h.). The difference spectrum below is with the INS spectrum of ice Ih as a standard reference (in orange and scaled by the right axis). As a result of an ordered H-bond network, amorphous Ih is characterized by a pronounced librational edge. Numbers and arrows highlight important modes summarized in Tables S1 and S2. The spectra were recorded at T < 20 K.

**Dried Ni-Hec before CO₂ Exposure.** In the spectrum for the dried Ni-Hec sample (Figure S), we observe peaks at 70, 109, 165, 215, and 300 cm⁻¹ (#1, #2, #3, #4, and #5), a broad distribution centered around 400 cm⁻¹ (#6), a broad distribution centered around 670 cm⁻¹ (#7), and a broad shoulder centered at 810 cm⁻¹ (#8). Previous INS measurements on β-Ni(OH)₂ have shown peaks at 90, 115, 294/313, 358, 394, 421/452, 680, and 880 cm⁻¹.⁴⁵,⁵₂ In agreement with the findings by Loch et al.⁴³ where the structure was identified as [Ni(OH)₀.₈₃(H₂O)₁.₁₇⁺]₁.₁₇⁻¹, the sharp peaks at 70 and 215 cm⁻¹ (#1, #4) are correlated with the confined water coordinated to the nickel hydroxide species.⁵³ One could question whether the peaks might be related to partial hydrolysis of Mg,F-moieties giving rise to a structure Mg₃OH. This is, however, unlikely as the Ni-Hec was obtained by ion exchange of melt-synthesized Na-Hec (Figure S3), which lacks bands in this region. Modes and tentative assignment are summarized in Table S1.

**Hydrated Ni-Hec before CO₂ Exposure.** For hydrated Ni-Hec equilibrated at 43% r.h., corresponding to a 2 WL hydration as in the PXRD/Raman spectroscopy measurements (Figure S), which contains both [Ni(H₂O)₆]²⁺ and [Ni(OH)₀.₈₃(H₂O)₁.₁₇]₀.₃₇¹.₁₇⁺, we observe a sharp peak on a broad distribution at 76 cm⁻¹ (#1), in addition to sharp peaks at 109, 215, and 237 cm⁻¹ (#2, #5, and #6), and minor vibrations centered around 130 and 165 cm⁻¹ (#3 and #4). From 335 to 1000 cm⁻¹, a broad distribution is found, with sharper features at 430, 480, 575, 650, and 800 cm⁻¹ (#8, #9, #11, #12, and #13). The hydrated Ni-Hec shows the same modes as the dried sample associated to the Ni–OH moieties as well as new modes associated to water (see further discussion below). Modes and the assignment are summarized in Table S2.

In the difference spectrum in Figure S, since we have removed the contribution from the clay and the condensed [Ni(OH)₀.₈₃(H₂O)₁.₁₇]₀.₃₇¹.₁₇⁺, we are left only with the contribution from different types of water present in this sample. Considering that INS is in particular sensitive to intramolecular water librational modes located between 300 and 1100 cm⁻¹, we can argue that the sharp librational edge at 336 cm⁻¹, previously observed in beidellite and montmorillonite⁵⁶ and shown to have a clear dependence on the interlayer cation, here originates from water coordinated to Ni⁴⁺. The second onset around 510 cm⁻¹ with a significant peak at 580 cm⁻¹ suggests the presence of less confined interlayer water, which is the noncoordinated interlayer water and responsible for the water loss below 100 °C observed in the TGA data (Figure S6). The absence of a clear librational edge similar to ice confirms the absence of a significant amount of noninterlayer water. The presence of these assigned water populations is further confirmed by analyzing the low energy region of the spectra. Free bulk water⁻¹ is characterized with an excitation around 50 cm⁻¹, while confined water molecules are characterized by a vibrational mode centered at about 80 cm⁻¹.⁵⁴,⁵⁵ Thus, the observation of the sharp peak at 76 cm⁻¹ is indicative of interlayer water coordinated to the Ni⁴⁺ cation, while the observation of a broad shoulder toward 50 cm⁻¹, also observed recently by Larsen et al.,⁵⁶ is consistent with the presence of water that is not coordinated with cations.

**Dried Ni-Hec after CO₂ Exposure.** Pronounced changes in the INS spectra are observed for the dried Ni-Hec (Figure 6a,b) upon CO₂ exposure. For the dried sample exposed to 5 bars of CO₂ at low energy transfer, a gain of intensity at 50 cm⁻¹, followed by an intensity loss at 72 cm⁻¹, and broad intensity gain between 80 and 160 cm⁻¹ and at 176 cm⁻¹ are observed. At 44 bar of CO₂, the signal in this region is more pronounced. Both at 5 and 44 bar, with the broad peaks centered around 215 and 400 cm⁻¹, a significant intensity loss was recorded. At 5 bar, there are still visible peaks left associated to the Ni–OH moieties at 336, 420, and 470 cm⁻¹, which are quite suppressed at 44 bar. Furthermore, at 5 bar, we observe an excess in energy at 675 cm⁻¹ that transforms into a sharp peak at 655 cm⁻¹ when increasing the CO₂ pressure to 44 bar.

For dried Ni-Hec after removing CO₂ by applying vacuum (Figure 6c), we observe at low wavenumbers a small decrease of intensity at 50, 70, and 101 cm⁻¹, whereas at higher energy transfers, the broad peaks centered at 215 and 400 cm⁻¹ have not reached their full initial prominence. The onset of the peak centered at 680 cm⁻¹ has a shift to 530 cm⁻¹ from 610 cm⁻¹. The signal seems to a large degree to return the initial signal, however, not to full prominence.
For dried Ni-Hec, the modes associated to the \([\text{Ni-(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17+}\) interlayer species are severely attenuated (Figure 6b) and considerable crystalline swelling due to CO\(_2\) exposure is observed (Figure 1). INS, gravimetric adsorption, and PXRD data on dried Na-Hec are presented in the Supporting Information, where neither uptake, swelling, nor spectroscopic changes are observed. As no changes are found when dried Na-Hec is exposed to CO\(_2\) and, in addition, previous observations on cations with high hydration enthalpy\(^8,10,22\) show no swelling, this suggest that the smectite-like interlayer does not play a role for the CO\(_2\) capture in the dried case. Thus, the \([\text{Ni-(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17+}\) species plays the important role in the adsorption of CO\(_2\) by dried Ni-Hec. This then means that the smectite layers remain unchanged with a d-spacing of about 10 Å and that the chlorite-like layers swell from about 13\(−\)14 to 15\(−\)16 Å. This reaction must be reversible since our PXRD experiments with repeated exposure to CO\(_2\), a long (>24 h) evacuation of the sample or moderate heating (>50 °C) is sufficient to return the Bragg reflection to the same initial position. This suggests that the observed state by INS after releasing the CO\(_2\) (Figure 6c) is a transient step during the return to the initial structure and that there is still some CO\(_2\) adsorbed to the structure. However, based on the present data, we cannot completely rule out the possibility that the reaction between CO\(_2\) and nickel hydroxide contained within the clay produces permanent structural changes to nickel hydroxide islands.

**Hydrated Ni-Hec after CO\(_2\) Exposure.** After being exposed to CO\(_2\) at room temperature and cooled down to \(T < 20 \text{ K}\), for the hydrated Ni-Hec (Figure 7), the changes are evident in the difference spectrum, and new modes are observed at 64, 95, and 117 cm\(^{-1}\). There is also a slight increase of intensity around 315 cm\(^{-1}\) and a small general loss of intensity between 400 and 900 cm\(^{-1}\). In addition, a new vibration is observed at 655 cm\(^{-1}\).

As various different hydrogenated species (coordinated and non-coordinated interlayer water, and \([\text{Ni-(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17+}\)) coexist for hydrated Ni-Hec, it is less obvious whether the changes observed by INS upon CO\(_2\) exposure are a consequence of changes in the librational modes associated to one or more of the water species or by direct association to the \([\text{Ni(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17+}\). The negligible changes along the librational edge agree with a strong coordination of water with Ni\(^{2+}\), which is not disrupted by CO\(_2\) exposure. However, we observe changes in the modes of the interlayer water by Raman spectroscopy (Figure 4), which together with the INS observations may be indicative of influence on the non-coordinated interlayer water by CO\(_2\). The general loss of INS intensity, similar to the dried case, suggests a similar interaction between CO\(_2\) and interlayer \([\text{Ni-(OH)}_{0.83}(\text{H}_2\text{O})_{1.17}]_{0.37}^{1.17+}\) both for dried and hydrated Ni-Hec. In addition, by Raman spectroscopy, we observe clear changes in what we identify as the OH-stretching modes of the intercalated nickel hydroxide, indicating an interaction with CO\(_2\). The smaller crystalline swelling of hydrated Ni-Hec compared to the dried case (Figure 2) is due to the layers being more similar to one another for hydrated Ni-Hec. We can assume that the smectite-like interlayer remains unchanged at about 15 Å, which is evidenced by the observed stability of the H-bonding by INS. Then, the chlorite-like layer swells from about 14−15 Å, and the chlorite-like layer could adopt a similar structure to what is observed for the dried state. The observed improved ordering is similar to what has previously associated...
been observed upon hydration of Na-Hec and suggests that the incorporation of CO2 in the interlayer strengthens the interaction between the interlayer species and 2:1 clay layers.

**Further Discussion.** For both the dried and the hydrated sample, a peak is forming at 655 cm$^{-1}$, (marked by an asterisk in Figure 6b), accompanied with vibrations at 64, 95, and 117 cm$^{-1}$. For the vibration at 655 cm$^{-1}$, we may suggest three possibilities based on the literature: Librational modes for water are observed in this region, bending modes for CO2 are also observed at 640 cm$^{-1}$ as well as modes associated to bicarbonates. As this peak is not observed in the spectra of the dried Na-Hec (see Supporting Information), formation of solid CO2 on the external surfaces of the Hec samples or in the remaining probing volume of the neutrons in the sample cell is unlikely. Notably the cross section of CO2 is 20 times lower than that of hydrogen; thus, the observation of this peak is unlikely to be due to the bending mode of CO2. It is possible that these changes also are associated with local changes of the water population. Following this reasoning, the peak that gains intensity at 64 cm$^{-1}$ could suggest a more ordered water population. The sharp vibration at 655 cm$^{-1}$ could also originate from a specific librational mode associated to this ordered water state, possibly through segregation of some water molecules to monomers and dimers. Water confinement in UO2F2 displays a peak at 778 cm$^{-1}$, that simulations attribute to a librational phonon mode along the a/b-axis at 712 cm$^{-1}$, due to confined water within the van der Waals gap of that material. The incorporation of CO2 in the interlayer may then force some of the H2O associated with the hydroxide species into a more confined situation. If we consider that the hydrogen bonds in HCO3$^-$ are weak, we may assign the present INS modes to HCO3$^-$ based on the study of Fillaux et al. on potassium bicarbonate. However, there are no observed bicarbonate modes by Raman spectroscopy (Figure 4).

From the current experiments, it still is unclear what is the reaction between CO2 and the [Ni(OH)$_{0.83}$(H$_2$O)$_{1.17}$]$_{0.37}^{1.17+}$ species. A possible mechanism that can be suggested has been observed on goethite under anhydrous conditions, where CO2 is chemisorbed according to Ni–OH + CO2 $\rightarrow$ Ni–O–(HCO3$^-$). By this reaction, hydroxide groups are converted into bicarbonate and this would significantly reduce the intensity associated to the corresponding modes. The chemisorbed CO2 may then be grafted as a bicarbonate to the condensed nickel hydroxide species, causing the swelling by adopting a structure as illustrated in Figure 8. However, as no bicarbonate modes are unambiguously observed by Raman spectroscopy, this is not evident from the present data.

Nevertheless, our experimental data show clear evidence that the intercalated [Ni(OH)$_{0.83}$(H$_2$O)$_{1.17}$]$_{0.37}^{1.17+}$ species play a crucial role in the incorporation of CO2 in the interlayers, both for the dried and hydrated cases. This is corroborated by DFT simulations (Figure 9), which are consistent with the experimentally observed $d$-spacing of the dehydrated system here and in the study by Loch et al. In addition, the simulations (Figure 9) show that the cohesion energy of the chlorite-like layers (−0.63 eV) is significantly lower than that for the smectite-like layers (−1.73 eV). This strengthens our arguments that the chlorite-like layers are responsible for the CO2 uptake and swelling in this system. Based on these calculations, we can argue that CO2 in the chlorite-like layers may attach to the edge of the nickel hydroxide islands and bind

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**Figure 8.** Hypothesized model for crystalline swelling in Ni-Hec clay in response to CO2 exposure. The CO2 interacts only with and causes swelling only of chlorite-like layers.

**Figure 9.** Basal spacing and energy per unit cell cohesive energy obtained by computational studies based on DFT calculations for the dehydrated smectite layers with Ni$^{2+}$ and chlorite-like layers with a confined β-Ni(OH)$_2$.
to not fully coordinated nickel atoms forming a strong bond ($E_{Ad}$), see Supporting Information.

**CONCLUSIONS**

Developing new sorbent materials for CO$_2$ is of key importance to solve issues related to greenhouse gas emissions. By PXRD, Raman spectroscopy, and INS, we identified a hitherto overlooked mechanism for CO$_2$ adsorption in hectorite clay. With these techniques, we took a new approach, where we study the mechanism for CO$_2$ adsorption from the perspective of interlayer hydrogenous species, that is hydroxides and water. Phase pure and charge homogeneous synthetic fluorohectorite provides a unique template for studying such interactions as it has no structural hydrogen incorporated in the 2:1 clay layer structure. For Ni-exchanged hectorite, an ordered interstratification of chlorite-like layers and smectite-like layers, we found that the intercalated nickel hydroxide promotes incorporation of CO$_2$ in the clay interlayers. Our results suggest that the CO$_2$ binds to the [Ni(OH)$_{0.83}$(H$_2$O)$_{1.17}$]$^{+0.37}$ species, both in the dried and the hydrated case, and is responsible for swelling in response to CO$_2$. No incorporation of CO$_2$ was found in the smectite-like layers for the dried case. Furthermore, CO$_2$ has limited impact on the water coordinated to the interlayer cations for the hydrated case.

In future work, we will perform ion exchange with a variety of transition metal elements known for their condensation perspective of interlayer hydronium species, that is fluorohectorite provides a unique template for studying such interactions as it has no structural hydrogen incorporated in the 2:1 clay layer structure. For Ni-exchanged hectorite, an ordered interstratification of chlorite-like layers and smectite-like layers, we found that the intercalated nickel hydroxide promotes incorporation of CO$_2$ in the clay interlayers. Our results suggest that the CO$_2$ binds to the [Ni(OH)$_{0.83}$(H$_2$O)$_{1.17}$]$^{+0.37}$ species, both in the dried and the hydrated case, and is responsible for swelling in response to CO$_2$. No incorporation of CO$_2$ was found in the smectite-like layers for the dried case. Furthermore, CO$_2$ has limited impact on the water coordinated to the interlayer cations for the hydrated case.

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**ASSOCIATED CONTENT**

Supporting Information

The Supporting Information is available free of charge at https://pubs.acs.org/doi/10.1021/acs.jpcc.0c07206.

Supporting INS and PXRD measurements of dried Na-Hec; diffractograms of hydrated Ni-Hec; gravimetric adsorption measurements on Ni-Hec and Na-Hec; thermogravimetric analysis of Ni-Hec and Na-Hec; and tables of assigned modes for INS (PDF)

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**Notes**

The authors declare no competing financial interest.

**ACKNOWLEDGMENTS**

This work was supported by the Research Council of Norway under the Frintake Program, project number 250728. This work was supported by the Deutsche Forschungsgemeinschaft (SFB 840). P.L. thanks Elite Network Bavaria in the framework of the Elite Study Program “Macromolecular Science” for support. A.K. and C.R.M. gratefully acknowledge the Research Centre for Gas Innovation (RCGI) and support from the Brazilian agencies FAPESP (2017/02317-9). This work was supported by the Research Council of Norway for support. A.K. and C.R.M. gratefully acknowledge the Research Centre for Gas Innovation (RCGI) and support from the Brazilian agencies FAPESP (2017/02317-9) and CNPq. We acknowledge the European Synchrotron Radiation Facility for provision of synchrotron radiation facilities and we would like to thank Wouter van Beek, Vadim Diadkin, and Dmitry Chernyshov for excellent assistance in using beamline SNBL BM01. Experiments at the ISIS Pulsed Neutron and Muon Source were supported by a beamtime allocation from the Science and Technology Facilities Council.64 Paulo Henrique Michels-Brito, Koki Araki, and Leander Michels are acknowledged for helpful discussions.

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