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Kinetic isotope effects of $^{12}$CH$_3$D + OH and $^{13}$CH$_3$D + OH from 278 to 313 K

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Abstract. Methane is the second most important long-lived greenhouse gas and plays a central role in the chemistry of the Earth’s atmosphere. Nonetheless there are significant uncertainties in its source budget. Analysis of the isotopic composition of atmospheric methane, including the doubly substituted species $^{13}$CH$_3$D, offers new insight into the methane budget as the sources and sinks have distinct isotopic signatures. The most important sink of atmospheric methane is oxidation by OH in the troposphere, which accounts for around 84% of all methane removal. Here we present experimentally derived methane + OH kinetic isotope effects and their temperature dependence over the range of 278 to 313 K for CH$_3$D and $^{13}$CH$_3$D; the latter is reported here for the first time. We find $k_{CH_4}/k_{CH_3}D = 1.31\pm0.01$ and $k_{CH_4}/k_{13CH_3}D = 1.34\pm0.03$ at room temperature, implying that the methane + OH kinetic isotope effect is multiplicative such that $(k_{CH_4}/k_{13CH_3}D)(k_{CH_4}/k_{CH_3}D) = k_{CH_4}/k_{13CH_3}D$, within the experimental uncertainty, given the literature value of $k_{CH_4}/k_{13CH_3} = 1.0039\pm0.0002$. In addition, the kinetic isotope effects were characterized using transition state theory with tunneling corrections. Good agreement between the experimental, quantum chemical, and available literature values was obtained. Based on the results we conclude that the OH reaction (the main sink of methane) at steady state can produce an atmospheric clumped isotope signal ($\Delta^{(13)}CH_3D = \ln(\left[CH_4\right][^{13}CH_3D]/[^{13}CH_4][CH_3D])$ of 0.02 $\pm$ 0.02. This implies that the bulk tropospheric $\Delta^{(13)}CH_3D$ reflects the source signal with relatively small adjustment due to the sink signal (i.e., mainly OH oxidation).

1 Introduction

Atmospheric methane is the subject of increasing interest from both the climate research community and the public due its impacts on climate change, as reported by the IPCC (2013). The direct radiative forcing of methane is 0.64 W m$^{-2}$. Including feedback mechanisms and secondary effects (e.g., increased O$_3$ production, stratospheric water vapor, and production of CO$_2$), methane’s radiative forcing becomes 0.97 W m$^{-2}$, two-thirds of the forcing by CO$_2$ over the same time period (IPCC, 2013, Fig. 8.15). Atmospheric methane has both natural and anthropogenic sources and the two categories contribute about equally (Ciais et al., 2013, and references therein). Wetlands are the dominant natural source, and agriculture and waste are the largest anthropogenic sources. Fossil fuels make smaller contributions. The majority (84%) of atmospheric methane is removed by oxidation by OH in the troposphere:

$$CH_4 + OH \rightarrow CH_3 + H_2O. \quad (R1)$$

while oxidation in the troposphere by Cl contributes about 4% of the total:

$$CH_4 + Cl \rightarrow CH_3 + HCl. \quad (R2)$$
About 8% of methane is removed in the stratosphere by radical oxidation, such as Reactions (R1), (R2) and (R3):

\[
\text{CH}_4 + \text{O}^{(1)} \rightarrow \text{CH}_3 + \text{OH}. \quad \text{(R3)}
\]

The rest (4%) is removed by soil (Kirschke et al., 2013).

Carbon and hydrogen isotopic analyses are widely used to distinguish microbial and thermal sources of atmospheric methane (e.g., Lowe et al., 1997; Ferretti et al., 2005; Tyler et al., 2007; Lassey et al., 2007). However, Reactions (R1), (R2), and (R3) produce relatively large D/H isotope effects (Saueressig et al., 1995, 1996, 2001; Crowley et al., 1999; Feilberg et al., 2005). Thus, the construction of an accurate top-down methane budget by isotopic analysis must take the kinetic isotope effect for the reaction 

\[
\text{CH}_3 + \text{OH} \rightarrow \text{CH}_2\text{D} + \text{H}_2\text{O},
\]

follows a stochastic distribution (i.e., \(\gamma\) for overcoming the problems of analysis of clumped CH
distributions (Ono et al., 2014):

\[
\Delta(\text{CH}_3\text{D}) \equiv \ln \left( \frac{[\text{CH}_3\text{D}][\text{H}_2\text{O}]}{[\text{CH}_2\text{D}][\text{H}_2\text{O}]} \right),
\]

where \([\text{CH}_3\text{D}]\), \([\text{H}_2\text{O}]\), \([\text{CH}_2\text{D}]\), and \([\text{H}_2\text{O}]\) represent the concentrations of the different isotopologues.

The kinetic isotope effects for the singly substituted species CH3D and CH4 have been studied previously both experimentally and theoretically; see Tables 3 and 4, respectively. The kinetic isotope effect for the reaction with OH is not described in the existing literature. The related kinetic isotope effect for the CH4 + Cl reaction was measured at room temperature with the present setup by Joelsson et al. (2014) and found to be 1.60 ± 0.04. The kinetic isotope effects in the doubly substituted methane isotopologue CH2D2 for the reaction CH4 + Cl has been studied previously by Feilberg et al. (2005).

In this work, the kinetic isotope effects \(D_\alpha\) and \(^{13}\text{C}\cdot\text{D}_\alpha\) are determined using the relative rate method. Species concentrations in the reaction cell are determined using Fourier transform infrared (FTIR) spectroscopy. Further, \(D_\alpha\), \(^{13}\text{C}_\alpha\), and \(^{13}\text{C}\cdot\text{D}_\alpha\) are calculated using quantum chemistry and transition state theory.

### 2 Experimental procedures

Sixteen experiments were conducted, numbered from 1 through 16 (see Table 1): eight (Experiments 1–8) for CH3D and eight (Experiments 9–16) for CH3D. The experiments were conducted at four different temperatures \((T = 278, 288, 298, 313\ K = 5, 15, 25, 40\ ^\circ\text{C})\); two experiments were conducted for each temperature.

#### 2.1 Relative rate method

The experiments were carried out using the relative rate method on a semi-static gas mixture. The decaying concentrations of reactants were measured as a function of the extent of reaction. Considering two bimolecular reactions with second-order rate coefficients \(k_A\) and \(k_B\).

\[
\text{A + OH} \rightarrow \text{products}, \quad (R4)
\]

\[
\text{B + OH} \rightarrow \text{products}, \quad (R5)
\]

and assuming there were no other loss processes, the following relation holds:

\[
\ln \left( \frac{[\text{A}_0]}{[\text{A}_t]} \right) = \frac{k_A}{k_B} \ln \left( \frac{[\text{B}_0]}{[\text{B}_t]} \right). \quad (4)
\]

Here \([\text{A}_0], [\text{A}_t], [\text{B}_0], \) and \([\text{B}_t]\) represent the concentrations of compounds A and B at times 0 and \(t\), respectively. The
Table 1. Experimental setup. The experiment numbers are listed in column Exp., the detector is listed in column Detect., the heavy CH₄ isotopologue included in the experiments is listed in column [¹³CH₃D], the mean measured temperatures in the photoreactor are listed in column T, the H₂O vapor concentrations at the start of the experiments (t = 0) as obtained from spectral fitting are listed in column [H₂O]₀, the mean O₃ concentrations after refill (i.e., the “top” values) as obtained from spectral fitting are listed in column [O₃]top, the ¹²CH₄ concentrations at the start of the experiments (t = 0) as obtained from spectral fitting are listed in column [¹²CH₄]₀, and the heavy CH₄ concentrations at the start of the experiments (t = 0) as obtained from spectral fitting are listed in column [¹³CH₃D]₀. Note that, for the experiment including CH₃D, the value of initial concentration only refers to [¹²CH₃D]₀.

<table>
<thead>
<tr>
<th>Exp.</th>
<th>Detect.</th>
<th>[¹³CH₃D]</th>
<th>T (K)</th>
<th>[H₂O]₀ (hPa)</th>
<th>[O₃]top (hPa)</th>
<th>[¹²CH₄]₀ (hPa)</th>
<th>[¹³CH₃D]₀ (hPa)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>MCT</td>
<td>CH₃D</td>
<td>298.2 ± 1.2</td>
<td>7.1</td>
<td>- a</td>
<td>0.030</td>
<td>0.054</td>
</tr>
<tr>
<td>2</td>
<td>MCT</td>
<td>CH₃D</td>
<td>297.6 ± 0.8</td>
<td>5.6</td>
<td>0.19</td>
<td>0.058</td>
<td>0.042</td>
</tr>
<tr>
<td>3</td>
<td>MCT</td>
<td>CH₃D</td>
<td>277.2 ± 0.2</td>
<td>5.2</td>
<td>0.29</td>
<td>0.109</td>
<td>0.046</td>
</tr>
<tr>
<td>4</td>
<td>MCT</td>
<td>CH₃D</td>
<td>277.0 ± 0.2</td>
<td>5.1</td>
<td>0.16</td>
<td>0.073</td>
<td>0.035</td>
</tr>
<tr>
<td>5</td>
<td>InSb</td>
<td>CH₃D</td>
<td>284.5 ± 0.1</td>
<td>7.2</td>
<td>0.26</td>
<td>0.025</td>
<td>0.033</td>
</tr>
<tr>
<td>6</td>
<td>InSb</td>
<td>CH₃D</td>
<td>291.1 ± 0.2</td>
<td>7.4</td>
<td>- a</td>
<td>0.052</td>
<td>0.050</td>
</tr>
<tr>
<td>7</td>
<td>InSb</td>
<td>CH₃D</td>
<td>313.5 ± 1.3</td>
<td>7.1</td>
<td>0.17</td>
<td>0.025</td>
<td>0.029</td>
</tr>
<tr>
<td>8</td>
<td>InSb</td>
<td>CH₃D</td>
<td>312.4 ± 0.9</td>
<td>4.3</td>
<td>- a</td>
<td>0.022</td>
<td>0.040</td>
</tr>
<tr>
<td>9</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>298.5 ± 0.1</td>
<td>5.1</td>
<td>- a</td>
<td>0.035</td>
<td>0.026</td>
</tr>
<tr>
<td>10</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>297.6 ± 0.6</td>
<td>6.4</td>
<td>0.13</td>
<td>0.025</td>
<td>0.033</td>
</tr>
<tr>
<td>11</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>276.8 ± 0.8</td>
<td>5.4</td>
<td>- a</td>
<td>0.024</td>
<td>0.024</td>
</tr>
<tr>
<td>12</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>277.2 ± 1.3</td>
<td>5.1</td>
<td>- a</td>
<td>0.022</td>
<td>0.030</td>
</tr>
<tr>
<td>13</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>287.4 ± 1.2</td>
<td>5.4</td>
<td>- a</td>
<td>0.021</td>
<td>0.028</td>
</tr>
<tr>
<td>14</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>287.4 ± 0.4</td>
<td>4.5</td>
<td>- a</td>
<td>0.016</td>
<td>0.029</td>
</tr>
<tr>
<td>15</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>314.4 ± 1.0</td>
<td>5.2</td>
<td>0.26</td>
<td>0.023</td>
<td>0.037</td>
</tr>
<tr>
<td>16</td>
<td>InSb</td>
<td>[¹³CH₃D]</td>
<td>313.8 ± 0.8</td>
<td>8.3</td>
<td>0.17</td>
<td>0.025</td>
<td>0.035</td>
</tr>
</tbody>
</table>

* Spectra recorded during or after photolysis, [O₃]top not available.

2.2 Photoreactor

Experiments were carried out in the photochemical reactor at the University of Copenhagen, Department of Chemistry. The reactor has been described in detail elsewhere (Nilsson et al., 2009). It consists of a 100 L quartz cell with multipass optics surrounded by 16 UV-C fluorescent Hg lamps in a temperature-controlled housing. The cell is coupled to a Bruker IFS 66v/S FTIR spectrometer with either a mercury cadmium telluride (MCT) (Experiments 1–4) or an indium antimonide (InSb) detector (Experiments 5–16). The InSb detector has a better signal-to-noise ratio; the MCT detector is used in Experiments 1–4 for logistical reasons. Two thermocouple gauges are placed inside the temperature-controlled housing to monitor the temperature and pressure; a gauge is connected to the cell to monitor pressure inside the cell. The temperatures and the pressures were logged every 0.5 s.

2.3 Laboratory procedure

Gas mixes were prepared by expanding H₂O vapor (Milli-Q ultrapure water) into the chamber through a glass gas manifold. The two methane isotopologues CH₃D (Experiments 1–8) (purity > 98 %, Cambridge Isotope Laboratories, Inc.) or [¹³CH₃D] (Experiments 9–16), and CH₄ (purity > 99 %, Aldrich) and O₃ were flushed into the chamber with a N₂ buffer (purity 99.998 %, Air Liquide), all at the concentrations given in Table 1. [¹³CH₃D] was synthesized using the Grignard reaction (see Joelsson et al., 2014). O₃ was generated from O₂ (purity 99.97 %, Air Liquide) using an ozone generator (model AC-20, O₃ Technology), preconcentrated before injection on silica gel cooled with ethanol and dry ice to −67 °C. The desired pressure in the cell (450 hPa) was obtained using N₂ as a bath gas. The starting pressure is chosen such that the pressure is high enough for the N₂ to quench O(¹D) radicals, but low enough to keep the final pressure below atmospheric pressure. The gas mixture was left to rest for up to 1.5 h while several IR spectra were recorded to ensure that no instability or dark chemistry occurs in the gas mix. The UV-C lamps were lit for up to 5 min photolyzing at least 75 % of the O₃ according to

O₃ + hν → O(¹D) + O₂. (R6)

O(¹D) then subsequently reacts with H₂O to yield OH:

O(¹D) + H₂O → OH + OH. (R7)

Up to 0.2 Pa of O₃ was flushed with about 20 hPa of N₂ into the chamber to compensate for the loss of ozone with time,
mainly due to $O^{(3)P} + O_3$. A pressure gradient was established and maintained throughout the filling process such that no gas leaked back from the chamber into the gas line. Spectra were recorded at each filling step. The procedure was repeated until the mix had a final pressure of 933 hPa. Two experiments were conducted at each of the temperatures 278, 288, 298, and 313 K for each of the two heavy methane isotopologues. Exact temperatures are listed in Table 1. After each experiment a dilution test was performed: 133 hPa was pumped out and the chamber is refilled with 133 hPa of $N_2$. This was repeated 5–6 times. Ideally, concentration calculations from the spectral fits (data analysis described below) of the resulting spectra should give a linear fit with the slope of 1. The slope of these dilution tests is presented in Table 2. In an extra experiment with $^{12}$CH$_3$D, $N_2$O (Air Liquide, no purity information available) was added as an O(1D) tracer. The results from this experiment are used as a benchmark to validate a model that was constructed to investigate the extent of O(1D) chemistry (see Sect. 2.5). An example of an experimental plot can be found in Fig. 4 and the full data set in Figs. S1–S8 in the Supplement.

2.4 Data analysis

The experimental IR spectra were analyzed using the program MALT, which simulates experimental FTIR spectra (Griffith, 1996) combined with nonlinear least-squares fitting to best-fit the calculated spectra to measured spectra (Griffith et al., 2012). The MALT program generates a simulated spectrum from initial estimates of the absorber concentrations and instrument parameters. The residual between the experimental and simulated spectra is reduced through iteration. Simulated line shapes are generated using HITRAN absorption parameters (version 2008) (Rothman et al., 2009) convolved with an FTIR instrument function simulating the Bruker IFS 66v/S instrument. The InSb detector covers a spectral range from 1800 to 5000 cm$^{-1}$ and the MCT detector covers a spectral range from 400 to 5000 cm$^{-1}$. The concentrations of $^{12}$CH$_3$D and $^{13}$CH$_3$D were calculated from spectral fits in the region 2140–2302 cm$^{-1}$ (see Figs. 1 and 2). Interference from $H_2$O, CO$_2$, and CO was eliminated by including simulated spectra obtained from the HITRAN database in the fit. As there are no HITRAN data available for $^{13}$CH$_3$D in this region, the cross sections from 2000 to 2400 cm$^{-1}$ for this isotopologue were estimated by shifting the spectrum of $^{12}$CH$_3$D (see Joelsson et al. (2014)). Concentrations of $^{12}$CH$_4$ were calculated from spectral fits in the region 2838–2997 cm$^{-1}$. Interference from $^{13}$CH$_3$D was reduced by including temperature-adjusted reference spectra in the fit, and interference from $^{12}$CH$_3$D, $H_2$O, and $H_2$CO was by including simulated spectra obtained from the HITRAN database in the fit (see Fig. 3). The spectral windows were sometimes adjusted to exclude saturated lines.

After $^4$CH$_3$D (where $x = 12$ or $x = 13$) and $^{12}$CH$_4$ were obtained from the spectral analysis, $\ln([^{13}$CH$_3$D]$_0/[^{13}$CH$_3$D]_t)$ was plotted against $\ln([^{12}$CH$_4]_0/[^{12}$CH$_4]_t)$ as described in Sect. 2.1. A straight line was fitted to these data points using a weighted total least-squares routine (York et al., 2004). The fitting procedure takes uncertainties in both dimensions into account. The uncertainty $\sigma (\ln([A]_0/[A]_t))$ was calculated
using standard error propagation:

$$\sigma(\ln([A]_0/[A]_t)) = \sqrt{\left(\frac{\sigma([A]_0)}{[A]_0}\right)^2 + \left(\frac{\sigma([A]_t)}{[A]_t}\right)^2},$$

where \(\sigma([A])\) was obtained as output from MALT. The fitting procedure, performed using a MATLAB script (York et al., 2004), also yields an error estimation which is defined as the uncertainty of the kinetic isotope effect \(\sigma(\alpha)\). An example of a straight line fit can be found in Fig. 4 and the full data set in Figs. S1–S8 in the Supplement. In the temperature dependence curve fitting procedure, the parameters \(A\) and \(B\) are from a linearized version of the Arrhenius equation:

$$\ln(k) = \ln(A) + B \cdot T^{-1},$$

which is fitted to the experimental data. Also here, the method of York et al. (2004) was used. The temperature in the cell was taken as the spatial average of the measurements from two thermocouples inside the temperature housing. The experiment temperature was defined by the temporal mean of the spatially averaged temperature measurement series and the uncertainty of the experiment temperature was the standard deviation of the spatially averaged temperature measurement series.

### 2.5 Kinetic model

A kinetic model was used to determine the influence of O(\(^{1}{D}\)), Reaction (R3), which rivals Reaction (R1). The model has been previously described and was used by Nilsson et al. (2012). However, only the methane reaction subset and associated (\(O_x\) and HO\(_x\)) chemistry were used here. The Kintecus program (Ianni, 2003) simulates the photolysis of O\(_3\) and the following oxidation chain of CH\(_4\). To model ozone photolysis accurately, the modeled O\(_3\) was matched to the measured value by adjusting the photolysis rate, then the model was verified by comparing the decrease in CH\(_4\) and the increase in H\(_2\)O during the experiment. The model was run for each refill of O\(_3\), where the reaction rates of Reactions (R1) and (R3) were obtained. The model was designed for room temperature; experiments at other temperatures were not modeled. Experiment 2 was thus modeled: 4.4% of CH\(_4\) was estimated to be lost to Reaction (R3). In an additional experiment, N\(_2\)O was introduced into the chamber as an O(\(^{1}{D}\)) tracer. Since N\(_2\)O does not react with OH and is not photolyzed at the wavelengths present (Nilsson et al., 2009), the decrease in N\(_2\)O should be only due to Reaction (R8):

$$\text{N}_2\text{O} + \text{O}(^{1}{D}) \rightarrow \text{products}.$$  \(\text{(R8)}\)

The amount of CH\(_4\) lost by Reaction (R3) can therefore be approximated by

$$1 - \frac{[\text{CH}_4]_t}{[\text{CH}_4]_0} = 1 - \exp\left(-k_{R3}[\text{O}(^{1}{D})]_t\right) = 1 - \frac{[\text{N}_2\text{O}]_0}{[\text{N}_2\text{O}]_t} \left(1 - \frac{k_{R8}}{k_{R3}}\right),$$

where \(k_{R8} = 1.27 \times 10^{-10} \text{ cm}^{-3} \text{s}^{-1}\) and \(k_{R3} = 1.75 \times 10^{-10} \text{ cm}^{-3} \text{s}^{-1}\) (Sander et al., 2011). This gives

Figure 3. A typical spectral fit in the region where \(^{12}\text{CH}_4\) is obtained (Experiment 5). Experimental data are shown by the topmost line, followed by the fitted partial (synthetic) spectra of the most dominant absorbers (CH\(_4\) and CH\(_3\)D), the resulting fitted (synthetic) spectrum (also including H\(_2\)O), and the residual between the measured and fitted spectra.

Figure 4. A typical experimental outcome, Experiment 5 and 6: CH\(_3\)D, \(T = 288\) K. Experimental data are shown using black open circles (Exp. 5) and black plus signs (Exp. 6). Corresponding dilution test points are shown using red symbols. A linear fit of the experimental points is shown using a black solid line, linear fits of the subsequent dilution test are represented by a red dotted line, and uncertainties for each point are represented by gray areas.
2.3 % [CH₄] lost by oxidation of O(¹D), Reaction (R3). The kinetic model described above estimated that 4.7 % [CH₄] was lost by Reaction (R3) for this additional experiment. Both methods agree that O(¹D) loss is a minor channel. No correction is applied, and the possible deviation is included in the estimated error.

3 Theoretical procedure

Rate constants and kinetic isotope effects for CH₄ + OH were calculated using a procedure similar to that employed by Joelsson et al. (2014).

3.1 Computational chemistry calculations

The geometries of reactants, products, and transition states were determined using a geometry optimization procedure based on the unrestricted MP2 method (Møller and Plesset, 1934) and the aug-cc-pVQZ orbital basis set (Dunning Jr., 1989; Woon and Dunning Jr., 1993). Harmonic vibrational frequencies for all relevant isotopologues of reactants, products, and transition states were obtained at the same level of theory. The calculations were carried out using the Gaussian 09 program package (Frisch et al., 2009).

The electronic energy of the optimized structures were refined using the CCSD(T) method (Watts et al., 1993; Knowles et al., 1993, 2000) with aug-cc-pVTZ, aug-cc-pVQZ, and aug-cc-pV5Z basis sets. The results from the different basis sets were used to extrapolate the electronic energy at the complete basis set limit following the approach of Halkier et al. (1998) as described by Joelsson et al. (2014).

3.2 Rate constant calculations

The abstraction of a CH₄ hydrogen atom by OH can occur at four different sites. Depending on the methane isotopologue in question, these sites are either distinguishable or indistinguishable. Microscopic rate constants are calculated for hydrogen abstraction at each site using classical transition state theory with a tunneling correction factor,

\[ k_{\text{micro}}^{(i)} = \frac{\eta_{\text{tun}}^{(i)} k_b T}{\hbar} \frac{Q_{\text{TS}}^{(i)}}{Q_{\text{reac}}} \exp(-\Delta E^{(i)}/RT), \]

where \( \Delta E^{(i)} \) is the reaction barrier height (including the zero-point vibrational energy) for the \( i \)th reaction path. \( \eta_{\text{tun}}^{(i)} \) is a tunneling correction factor obtained using the Wigner tunneling correction (Wigner, 1932). \( Q_{\text{TS}}^{(i)} \) and \( Q_{\text{reac}} \) are the partition functions for the transition state and reacting pair, respectively. The total rate constant is obtained by summing over the microscopic rate constants.

4 Results and discussion

The results of the 16 individual experiments (eight for each isotopologue) are tabulated in Table 2. The resulting \( D\alpha, ^{13}C\alpha, \) and \( ^{13}C.D\alpha \) values from the experimental and theoretical studies are tabulated in Tables 3, 4, and 5 along with previous experimental and theoretical results. The results are also shown in Figs. 5 and 6 for reactions of CH₃D and \(^{13}\text{CH}_3\)D, respectively.

The exponential curve fits yielded the parameters presented in Tables 3 and 5 giving \( D\alpha_{\text{Arr}} = 1.32 \pm 0.13 \) and \( ^{13}C.D\alpha_{\text{Arr}} = 1.32 \pm 0.20 \) at \( T = 298 \text{ K} \). The mean of results of room temperature experiments Experiments 1 and 2 and Experiments 9 and 10 is \( D\alpha_{\exp} = 1.31 \pm 0.01 \) and \( ^{13}C.D\alpha_{\exp} = 1.34 \pm 0.03 \), respectively. It follows that \( y_{\exp} = 1.02 \pm 0.02 \) at \( T = 298 \text{ K} \), using \( ^{13}C.D\alpha_{\exp} = 1.0039 \pm 0.0002 \) (Saueressig et al., 2001), meaning that the clumped isotope might react slower relative to what would be predicted based on the kinetic isotope effects of CH₃D and \(^{13}\text{CH}_3\)D. However, if the Arrhenius parameters are used to calculate the kinetic isotope effects at \( T = 298 \text{ K} \), \( y_{\text{Arr}} = 1.00 \pm 0.18 \) (i.e., the reaction has no clumping effect). All uncertainties are given as 1 standard deviation (σ). The theoretical results give \( y_{\text{theory}} = 1.00 \) at \( T = 298 \text{ K} \).

The present experimental room temperature results for \( D\alpha \) agree, to within the error bars, with the previous experimental studies, with the exception of DeMore (1993). DeMore (1993) used FTIR spectroscopy with a slow flow setup where the two methane isotopologues were measured separately with a common reference compound. The low \( D\alpha \) value in DeMore (1993) may be explained by interference from...
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Table 3. Experimental and theoretical studies of $^{D} \alpha$. The temperature dependence studies are presented in the Arrhenius form $^{D} \alpha(T) = A(T/298 \text{K})^n \exp(BT^{-1})$, where $A$, $n$, and $B$ are tabulated; $T$ is the given temperature range; and $^{D} \alpha(T = 298 \text{K})$ is the resulting kinetic isotope effect at $T = 298 \text{K}$. Where no $B$ coefficient is presented, $A$ can be taken as $^{D} \alpha(T)$ for the given temperature range. All uncertainties are given as 1 standard deviation ($\sigma$).

<table>
<thead>
<tr>
<th>Study</th>
<th>$A$</th>
<th>$n$</th>
<th>$B$ (K)</th>
<th>$T$ (K)</th>
<th>$^{D} \alpha(T = 298 \text{K})$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Experimental studies</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Present study$^a$</td>
<td>1.23 ± 0.08</td>
<td>0</td>
<td>21 ± 21</td>
<td>[278, 313]</td>
<td>1.31 ± 0.01$^h$</td>
</tr>
<tr>
<td>Saueressig et al. (2001) $^b$</td>
<td>1.294 ± 0.009</td>
<td>–</td>
<td>–</td>
<td>296</td>
<td>–</td>
</tr>
<tr>
<td>Gierczak et al. (1997) $^c$</td>
<td>1.09 ± 0.05</td>
<td>0</td>
<td>49 ± 11</td>
<td>[220, 415]</td>
<td>1.25 ± 0.07</td>
</tr>
<tr>
<td>DeMore (1993) $^a$</td>
<td>0.91</td>
<td>0</td>
<td>75</td>
<td>[298, 360]</td>
<td>1.17</td>
</tr>
<tr>
<td>Gordon and Mulac (1975) $^d$</td>
<td>1.50</td>
<td>–</td>
<td>–</td>
<td>416</td>
<td>–</td>
</tr>
<tr>
<td>Gordon and Mulac (1975) $^{d,i}$</td>
<td>1.06</td>
<td>–</td>
<td>–</td>
<td>416</td>
<td>–</td>
</tr>
<tr>
<td>Theoretical studies</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Present study$^e$</td>
<td>1.314</td>
<td>0</td>
<td>6.354</td>
<td>[200, 300]</td>
<td>1.339</td>
</tr>
<tr>
<td>Selleväg et al. (2006) $^f$</td>
<td>[1.30, 1.00]</td>
<td>–</td>
<td>–</td>
<td>[200, 1500]</td>
<td>1.27</td>
</tr>
<tr>
<td>Masgrau et al. (2001) $^{g,j}$</td>
<td>1.00</td>
<td>–0.02</td>
<td>50.5</td>
<td>[200, 1500]</td>
<td>1.25</td>
</tr>
</tbody>
</table>


Table 4. Experimental and theoretical studies of $^{13}$C$\alpha$. The temperature dependence studies are presented in the Arrhenius form $^{13}$C$\alpha(T) = A \exp(BT^{-1})$, where $A$ and $B$ are tabulated, $T$ is the given temperature range, and $^{13}$C$\alpha(T = 298 \text{K})$ is the resulting kinetic isotope effect at $T = 298 \text{K}$. Where no $B$ coefficient is presented, $A$ can be taken as $^{13}$C$\alpha(T)$ for the given temperature range. All uncertainties are given as 1 standard deviation ($\sigma$).

<table>
<thead>
<tr>
<th>Study</th>
<th>$A$</th>
<th>$B$ (K)</th>
<th>$T$ (K)</th>
<th>$^{13}$C$\alpha(T = 298 \text{K})$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Experimental studies</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Saueressig et al. (2001) $^a$</td>
<td>1.0039 ± 0.0002</td>
<td>–</td>
<td>296</td>
<td>–</td>
</tr>
<tr>
<td>Cantrell et al. (1990) $^b$</td>
<td>1.0054 ± 0.0005</td>
<td>–</td>
<td>–</td>
<td>[200, 1500]</td>
</tr>
<tr>
<td>Rust and Stevens (1980) $^{c,i}$</td>
<td>1.003</td>
<td>–</td>
<td>–</td>
<td>–</td>
</tr>
<tr>
<td>Theoretical studies</td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>Present study$^d$</td>
<td>1.0137</td>
<td>–1.219</td>
<td>[200, 300]</td>
<td>1.0094</td>
</tr>
<tr>
<td>Selleväg et al. (2006) $^e$</td>
<td>[1.014, 1.00]</td>
<td>–</td>
<td>[200, 1500]</td>
<td>1.003</td>
</tr>
<tr>
<td>Gupta et al. (1997) $^f$</td>
<td>1.010</td>
<td>–</td>
<td>300</td>
<td>–</td>
</tr>
<tr>
<td>Melissas and Truhlar (1993) $^{g,h}$</td>
<td>[1.005, 1.001]</td>
<td>–</td>
<td>[223, 416]</td>
<td>1.005</td>
</tr>
</tbody>
</table>

$^a$ Tunable diode laser absorption spectroscopy–isotope ratio mass spectrometry. $^b$ Gas chromatography–mass spectrometry. $^c$ Isotope ratio mass spectrometry. $^d$ Transition state theory with Wigner tunneling correction. $^e$ Canonical variational transition state theory. $^f$ Conventional transition state theory. $^g$ Interpolated variational transition state theory with centrifugal-dominant, small-curvature tunneling coefficients. $^h$ Ab initio calculations. $^i$ No temperature information available.

O($^1$D) radicals: OH radicals were produced by photolysis of O$_3$ at 254 nm in the presence of H$_2$O, as in the present study. A relatively high rate of Reaction (R3) would reduce the final kinetic isotope effect, since the kinetic isotope effect for oxidation with O($^1$D) is smaller than the kinetic isotope effect for the oxidation of OH (Saueressig et al., 2001). The present experimental results of $^{D} \alpha$ are also in good agreement with Masgrau et al. (2001) and Selleväg et al. (2006). The theoretical calculations of $^{D} \alpha$ give a slightly higher value than the experimental results, although they are in good agreement with the best Arrhenius curve fit at $T = 298 \text{K}$.

The experimental room temperature result for $^{13}$C$^{D} \alpha = 1.34 ± 0.03$ agrees, to within the error bar, with both the theoretical value and the best estimate of the Arrhenius curve fit at $T = 298 \text{K}$.

The theoretical calculations show a very small temperature dependence; the variability in the experimental data is large compared to the value of the slope, making a quantification of the temperature dependence uncertain. Furthermore, the
Table 5. Experimental and theoretical studies of $^{13}$C,D$\alpha$. The temperature dependencies are presented in the Arrhenius form $^{13}$C,D$\alpha(T) = A \exp(BT^{-1})$, where $A$ and $B$ are tabulated, $T$ is the given temperature range, and $^{13}$C,D$\alpha(T = 298 \text{ K})$ is the resulting kinetic isotope effect at $T = 298 \text{ K}$. All uncertainties are given as 1 standard deviation ($\sigma$).

<table>
<thead>
<tr>
<th>Study</th>
<th>$A$</th>
<th>$B$ (K)</th>
<th>$T$ (K)</th>
<th>$^{13}$C,D$\alpha(T = 298 \text{ K})$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Experimental study</td>
<td>1.18 ± 0.10</td>
<td>38 ± 26</td>
<td>[278, 313]</td>
<td>1.34 ± 0.03$^c$</td>
</tr>
<tr>
<td>Theoretical study</td>
<td>1.328</td>
<td>5.301</td>
<td>[200, 300]</td>
<td>1.349</td>
</tr>
</tbody>
</table>

$^a$ Fourier transform infrared spectroscopy – relative rate. $^b$ Transition state theory with Wigner tunneling correction. $^c$ Average room temperature value, not obtained from curve fit.

Figure 5. The individual measurements of D$\alpha$ are represented by black points, the accompanying individual error bars are represented by thin black solid lines, the experimental Arrhenius curve fit is represented by a black dashed line, the theoretical Arrhenius curve fit is represented by a blue dashed-dotted line, the mean of the room temperature measurements is represented by a red solid circle, and the uncertainty of the room temperature measurements is represented by a thick red solid line. All uncertainties are given as 1 standard deviation ($\sigma$).

Figure 6. The individual measurements of $^{13}$C,D$\alpha$ are represented by black points, the accompanying individual error bars are represented by thin black solid lines, the experimental Arrhenius curve fit is represented by a black dashed line, the theoretical Arrhenius curve fit is represented by a blue dashed-dotted line, the mean of the room temperature measurements is represented by a red solid circle, and the uncertainty of the room temperature measurements is represented by a thick red solid line. All uncertainties are given as 1 standard deviation ($\sigma$).

Atmospheric implications

At steady state, assuming no clumping in emissions, $\Delta(^{13}\text{CH}_3\text{D}) = \ln(\gamma)$. It follows that $\Delta(^{13}\text{CH}_3\text{D}) = 0.02 \pm 0.02$, implying that the clumped isotope effect of the OH reaction is small. Wang et al. (2015) studied a variety of natural samples and found $\Delta(^{13}\text{CH}_3\text{D})$ values that vary from $(-3.36 \pm 1.42)\%e$ for samples of bubbling gases from The...
Cedars, Cazadero, California, USA, to (6.17 ± 0.34)%e for samples from gas voids and hydrates in sediment cores from the northern Cascadia margin. Methanogenic bacteria in the laboratory ranged from −1.34 to 2.29%e. Natural gas was found to vary from 2.53 to 5.18%e. Given these natural variations in $\Delta^{13}$(CH3D) values of the sources, it is clear from isotopic mass balance that the fractionation from atmospheric oxidation by OH will likely have little effect on the observed $\Delta^{13}$(CH3D) value of tropospheric methane. Rather, this number will reflect the relative contributions of the sources. $\Delta^{13}$(CH3D) would therefore be a more straightforward tracer for tracking methane sources than conventional isotopic analysis. Further field studies are clearly needed to characterize $\Delta^{13}$(CH3D) values of microbial and thermal methane.

5 Conclusions

We present experimentally derived CH4 + OH kinetic isotope effects and their temperature dependence for CH3D and $^{13}$CH3D; the latter is reported for the first time. We find $D_\alpha = 1.31 ± 0.01$ and $^{13}$CD $\alpha = 1.34 ± 0.03$ at room temperature, implying that the kinetic isotope effect is multiplicative such that $(k_{CH4}/k_{CH3D})(k_{CH4}/k_{CH3D}) = k_{CH4}/k_{CH3D}$ to within the experimental uncertainty. We compare our experimental results to theoretical estimates derived using transition state theory with tunneling correction and kinetic isotope effects reported in the literature. We find good agreement between theoretical and literature values. Based on these experiments we find that the OH reaction (the main sink of methane) at steady state has a clumped $\Delta^{13}$(CH3D) = 0.02 ± 0.02.

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References


