One-dimensional coordination polymers of [Co3(dpa)4]2+ and [MF6]2- (M = ReIV, ZrIV and SnIV)

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One-dimensional coordination polymers of alternating metal–metal bonded trinuclear \([\text{Co}_3\text{dpa}_4]^{2+}\) and \([\text{MF}_6]^{2-}\) (\(M = \text{Re}^{IV}, \text{Zr}^{IV}\) and \(\text{Sn}^{IV}\))

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One-dimensional coordination polymers of alternating metal–metal bonded trinuclear \([\text{Co}_3\text{dpa}_4]^{2+}\) (dpa = the anion of 2,2′-dipyridylamine) building blocks and \([\text{ReF}_6]^{2-}\) (1), \([\text{ZrF}_6]^{2-}\) (2) or \([\text{SnF}_6]^{2-}\) (3) linkers have been self-assembled and crystallographically characterized. Magnetic measurements reveal a significant ferromagnetic coupling \((J/K = 9.9\text{K})\) between \(S = 1/2\) \((\text{Co}^{II}_3)\) and \(S = 3/2\) \(\text{Re}^{IV}\) magnetic sites through a single, unsupported fluoride bridge in 1.

One-dimensional metal–ligand coordination polymers are of particular interest in the molecular magnetism community for their potential as single-chain magnets (SCMs). Because the stereochemistry around the metal ion determines the topology of coordination-driven self-assembled systems, a simple approach to magnetic linear polymers entails the use of paramagnetic metal complexes possessing two empty, or at least kinetically labile, coordination sites. These acceptors can then be associated into chains using paramagnetic linkers with two donor sites, the most common of which have been \(\text{trans}\)-cyanidoetalatellates. Possessing rich magnetic and electronic properties, in addition to axial geometry, paddlewheel complexes are attractive building blocks for the construction of linear one-dimensional systems. Indeed, many dinuclear metal ion tetracarboxylates, e.g. \({\{\text{Cr}_2{\text{O}}\}}^4-\), \({\{\text{Cu}_2{\text{O}}\}}^4-\), \({\{\text{Rh}_2{\text{O}}\}}^4-\) and \({\{\text{Ru}_2{\text{O}}\}}^4-\), have strongly Lewis acidic axial sites and form extended chain structures in the absence of exogenous ligands by coordination of a carboxylate oxygen atom to the axial position of an adjacent molecule. On the other hand, copper and zinc paddlewheels have been extensively used in the formation of three-dimensional metal–organic frameworks, while one-dimensional coordination polymers using exogenous spacers and \({\{\text{Rh}_2{\text{O}}\}}^4-\), \({\{\text{Cu}_2{\text{O}}\}}^4-\), \({\{\text{Re}_2{\text{O}}\}}^4-\), \({\{\text{Zn}_2{\text{O}}\}}^4-\) units are also well represented. Of particular note are magnetic systems, some quite sophisticated, constructed of \({\{\text{Ru}_2{\text{O}}\}}^5-\) \((S = 3/2)^9\) and \({\{\text{Ru}_2{\text{O}}\}}^4-\) \((S = 1)^{10}\) building blocks and which display remarkable properties.

We here turn our attention to Extended Metal Atom Chains (EMACs), which contain three or more metal ions typically supported by oligopyridylamines. While dinuclear divalent paddlewheel complexes are often diamagnetic, EMACs tend to have an odd number of metal centers and to be paramagnetic. This is due to a variety of factors, including an odd number of electrons, accidental degeneracy of frontier orbitals, or disymmetry in the linear complex giving rise to an isolated high-spin metal ion. However, with the exception of a few coordination polymers based on \([\text{Ni}_3\text{dpa}_4]{\text{Cl}_2}\)^1, EMACs have rarely been used to build extended systems.

The \([\text{Co}_3\text{dpa}_4]{\text{Cl}_2}\)^2+ unit (Chart 1) is expected to be a useful building block, based on the interesting physical properties of the \([\text{Co}_3\text{dpa}_4]{\text{Cl}_2}\) analogue. Its core consists of three aligned cobalt metal ions possessing a delocalized 3-electron 3-center bond, and depending on crystal packing effects, the spacing of the Co ions can be equal (as in the orthorhombic \([\text{Co}_3\text{dpa}_4]{\text{Cl}_2}\), \(\text{CH}_2\text{Cl}_2\) phase), or unequal, (as in the tetragonal \([\text{Co}_3\text{dpa}_4]{\text{Cl}_2}\), \(\text{CH}_2\text{Cl}_2\) phase). Both forms show a spin-crossover (SCO) process from \(S = 1/2\) to \(S = 3/2\) or 5/2. Remarkably, the one electron oxidized compound, \(\text{Co}_3\text{dpa}_4\text{Cl}_2\text{[BF}_4]\) undergoes a two-step SCO, from \(S = 0\) to \(S = 1\) to \(S = 2\). Finally, the axial chloride ligands can be

![Chart 1](https://www.rsc.org/chemcomm)

Schematic representation of the \([\text{Co}_3\text{dpa}_4]{\text{Cl}_2}\)^2+ unit.

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† Electronic supplementary information (ESI) available: Physical methods, IR, TGA, PXRD, magnetization data, crystal structure data for 1–3 and \((\text{PPh}_4)\text{[ReF}_6]\) MeCN. CCDC 1415823–1415826. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5cc06704a
cleanly removed from [Co6(dpa)4Cl2] using silver salts, generating MeCN,15 BF4−,16 CN−, N(CN)2− or NCS− adducts. Nonetheless, [Co6(dpa)4]2+ has not been widely used in extended structures, even if organization into rigid multi-dimensional systems is an appealing strategy to introduce cooperativity in the SCO process and thus to generate possible spin transition phenomena. Recent work by Shatruk and coworkers on putative 2D grids formed by four ditopic [Co6(dpa)4]2+ units coordinated to [Co(CN)6]3− or [Fe(CN)6]3− is the only example of the assembly of [Co6(dpa)4]2+ units into extended networks.18 Unfortunately, these materials could not be crystallographically characterized.

In order to obtain 1 : 1 neutral chains, we have selected the dianionic metalloligands [ReF6]2−, [ZrF6]2− and [SnF6]2− to link the [Co6(dpa)4]2+ building block. (PPh4)2[ReF6]2H2O was recently reported to exhibit a pronounced magnetic anisotropy,19 a property that was retained in 1D coordination polymers featuring the [Co3(dpa)4]2+ building block. (PPh4)2[ReF6] immediately gave a dark red solution of [Co3(dpa)4(BF4)2] in a thin tube and left to slowly diffuse over several weeks.‡ Green blocks of [Co3(dpa)4(ReF6)] 2DMF (1), [Co3(dpa)4(ZrF6)] 2DMF (2) and [Co3(dpa)4(SnF6)] 2DMF (3) were obtained in moderate yield. Powder X-ray diffraction revealed the presence of only one phase and thermal gravimetric analysis was consistent with the presence of two DMF molecules (ESI†).

Compounds 1–3 are isostructural. They crystallize in the space group P4/ncc (Table S1, ESI†) with the four-fold axis coincident with the rigorously linear Co–F–M axis. The chains are racemic, being made up of alternating A and A [Co6(dpa)4]2+ helicoidal moieties (Fig. 1). A small disorder is found with respect to the wrapping of the dpa− ligands, and each position is occupied by ca. 80% of one enantiomer and 20% of the other. The [Co6]2+ core is slightly asymmetrical with differences in Co–Co distances of 0.008(1), 0.013(1) and 0.013(1) Å in 1, 2 and 3, respectively (at 200 K; Table 1). Surprisingly, the Co–F distances are quite unequal, with differences of 0.066(4) Å (1), 0.059(4) Å (2) and 0.069(4) Å (3), with the longer distance associated with the terminal cobalt engaged in the longer Co–Co bonding interaction. This asymmetry is however not reflected in the two individual M–Fax distances, which are distinctly and equally elongated due to their coordination to the Co centers (Table 1).

Magnetic susceptibility measurements were performed on polycrystalline samples of 1–3 between 1.85 and 300 K. At room temperature, the χT product of 1 amounts to 2.4 cm³ K mol⁻¹, slightly higher than the theoretical value of 1.97 cm³ K mol⁻¹ for isolated 1 = ½ [Co6(dpa)4]2+, g = 2.35, C = 0.52 cm³ K mol⁻¹) and

<table>
<thead>
<tr>
<th>Table 1</th>
<th>Selected bond distances (Å)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>1</td>
</tr>
<tr>
<td>Co1–Co3</td>
<td>2.270(1)</td>
</tr>
<tr>
<td>Co2–Co3</td>
<td>2.278(1)</td>
</tr>
<tr>
<td>Co1–F1</td>
<td>2.115(4)</td>
</tr>
<tr>
<td>Co2–F2</td>
<td>2.181(4)</td>
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<tr>
<td>M–F1</td>
<td>1.962(4)</td>
</tr>
<tr>
<td>M–F2</td>
<td>1.970(4)</td>
</tr>
<tr>
<td>M–Feq</td>
<td>1.931(3)</td>
</tr>
</tbody>
</table>

S = 3/2 ([ReF6]2−, g = 1.76, C = 1.45 cm³ K mol⁻¹) spins (Fig. 2). On lowering the temperature, the χT product increases steadily, with a more abrupt increase below ca. 30 K eventually reaching 9.6 cm³ K mol⁻¹ at 1.85 K, suggestive of a significant ferromagnetic coupling between [Co6]2+ and ReIV magnetic sites. The data were fit to a Seiden model20 derived from the exchange-coupling Hamiltonian $\hat{H} = -2J \sum_{\text{pairs}} (S_i \cdot (\hat{S}_i + \hat{S}_{i+1}))$, where $\hat{S}_i$ and $\hat{S}_{i+1}$ represent spin-operators of [Co6]2+ and ReIV, yielding g = 2.05(3) and $J/K$ = +9.9(1) K. As expected, this average g value falls in between the previously reported values for [Co6(dpa)4Cl2] 2CH2Cl2 (g = 2.35)13 and [PPh4][ReF6] 2H2O (g = 1.76).19 Compounds 2 and 3, possessing diamagnetic [ZrF6]2− and [SnF6]2− linkers, show an almost temperature independent χT product between about 50 and 300 K. The low temperature decrease of the χT product is likely due to weak antiferromagnetic coupling between the...
of CO$_{3}^{3+}$) centers through the diamagnetic linkers, and fitting the data to a regular quantum $s = 1/2$ spin chain model$^{21}$ derived from the Hamiltonian $H = -2J \sum_{i=1}^{N} \delta_{i} \cdot \delta_{i+1}$ yields $J/k_{B} = -1.0$ K for 2 while magnetic interactions are virtually undetectable above 1.85 K for 3 (with $g = 2.36$ for both). The difference in the magnetic coupling between 2 and 3 may be attributed to the lack of empty d orbitals in Sn$^{4+}$ to mediate a superexchange interaction, in contrast to Zr$^{4+}$ in 2.

The magnetic behaviour of these coordination polymers is significantly different to that of their building units. For example, the parent compound [Co$_{3}$(dpa)$_{4}$Cl$_{2}$].CH$_2$Cl$_{2}$ demonstrates a SCO event above 200 K.$^{13}$ No such thermal behaviour is observed in 1 and 3 up to 300 K, while for 2, an onset of what may be a SCO event appears reversibly above 250 K. However, above 300 K, a rapid and irreversible rise in the $\gamma$ powder pattern of these compounds was observed in concert with DFM loss. Therefore, the assembly of ostensibly SCO [Co$_{3}$(dpa)$_{4}$]$^{2+}$ units with [MFe]$^{2+}$ linkers into chains increased the SCO temperature to significantly different to that of their building units. For example, the parent compound [Co$_{3}$(dpa)$_{4}$Cl$_{2}$].CH$_2$Cl$_{2}$ demonstrates a SCO event above 200 K.$^{13}$ No such thermal behaviour is retained upon assembly of [ReF$_{6}$]$^{2-}$ from the Hamiltonian $H = -2J \sum_{i=1}^{N} \delta_{i} \cdot \delta_{i+1}$ yields $J/k_{B} = -1.0$ K for 2 while magnetic interactions are virtually undetectable above 1.85 K for 3 (with $g = 2.36$ for both). The difference in the magnetic coupling between 2 and 3 may be attributed to the lack of empty d orbitals in Sn$^{4+}$ to mediate a superexchange interaction, in contrast to Zr$^{4+}$ in 2.

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**Notes and references**

† General synthesis of [Co$_{3}$(dpa)$_{4}$][MFe]$^{2+}$:DMF, 1–3: [Co$_{3}$(dpa)$_{4}$][MFe]$^{2+}$ (50 mg, 0.05 mmol) and AgBF$_{4}$ (20 mg, 0.10 mmol) were combined in 10 mL of DMF in a glovebox. The mixture was stirred overnight and filtered. Anhydrous [PPh$_{4}$][MFe]$^{2+}$ (1 eq.) was dissolved in 10 mL of MeCN. The MeCN solution was layered upon the DMF solution, separated by a 1:1 mixture of DMF:MeCN in glass tubes (200 mm, $\odot$ 10 mm). Dark green blocks were collected from the walls of the tubes after 4 weeks.


